

## Nano-porous Adsorbent from Fruit Peel Waste for Decolorization Studies

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### ABSTRACT

The aim of the present study was to discover the effect of low-cost and ecofriendly nano-porous adsorbent (orange peel) as an ideal alternative to the current expensive methods of removing dye Bromophenol blue (Bb) from aqueous solution. The effects of various experimental parameters (e.g., dye concentration, temperature and pH) were investigated by batch equilibrium experiments. The fitness of Langmuir model in the present system shows the formation of monolayer coverage of the adsorbate at the outer space of the adsorbent. Freundlich model isotherm was analyzed. The monolayer adsorption capacity determined was reasonably high ( $\text{mg g}^{-1}$ ) at Temperature (30, 45, 60°C) 88.11 ( $\text{mg g}^{-1}$ ) and pH (5.5, 6.8, 7.8) 149.760 ( $\text{mg g}^{-1}$ ) for adsorption of Bromophenol blue dye, respectively. The values of dimensionless equilibrium parameter like separation factor ( $R_L$ ) at different particle size and temperature indicates the favorability of the process described in the present study. The kinetics of Bromophenol blue dye adsorption nicely followed pseudo-first and second order rate expression and a significant role in the adsorption mechanism. Langmuir and Freundlich models could be used to describe dye sorption on nano porous adsorbent at equilibrium and also the kinetic data gave a better fit. The model fits the second order equation well with correlation coefficients of 0.99967. The adsorption behavior of Bb on orange peel was investigated by FT-IR technique. We concluded on the basis of the results shown in this study that orange peel was an attractive candidate for removing Bb dye from the dye wastewater.

**Key words:** Orange peel waste, bromophenol blue, decolorization, adsorption isotherms, wastewater treatment

### INTRODUCTION

Many industries which used dyes and pigments generated wastewater, characteristically high in color and organic content. Presently, it was estimated about 10,000 of different commercial dyes and pigments exists and over  $7 \times 10^5$  tones are produced annually world wide (Grag *et al.*, 2004). Dyes are widely used in industries such as textile, rubber, paper, plastic, cosmetic etc. Among these various industries, textile ranks first in usage of dyes for coloration of fiber (Azahar *et al.*, 2005; Grag *et al.*, 2004).

Many techniques have been found for removal of dye-containing wastewater such as chemical oxidation, membrane filtration, biodegradability, separation and adsorption techniques (Koyuncu, 2009; Choy *et al.*, 1999). Among these, the adsorption process is one of the effective methods for

removal dyes from the waste effluent. The process of adsorption has an edge over the other methods due to its sludge free clean operation and completely removed dyes, even from the diluted solution. Activated carbon (powdered or granular) is the most widely used adsorbents because it has excellent adsorption efficiency for the organic compound. But, commercially available activated carbon is very expensive. Furthermore, regeneration using solution produced small additional effluent while regeneration by refractory technique results in a 10-15% loss of adsorbents and its uptake capacity (Azahar *et al.*, 2005; Wang *et al.*, 2005).

Nowadays, there is numerous numbers of low cost, commercially available adsorbents coal, fly ash, wood, silica gel, clay materials (bentonite, montmorillonite, etc.), agricultural wastes (bagasse pith, maize cob, coconut shell, rice husk, etc.) and cotton wastes and cellulose based wastes such as orange, lemon, banana and lychee (Bhatnagar and Minocha, 2010; Thirumavalavan *et al.*, 2010; Bhatnagar *et al.*, 2010; Memon *et al.*, 2009; Achak *et al.*, 2009) which had been used for the dye removal (Azahar *et al.*, 2005; Annadurai *et al.*, 2002; Juang *et al.*, 1997; Theng and Wells, 1995; Singh and Rawat, 1994). Thereby, the present investigation was made on cellulose-based wastes of *Citrus reticulata* for adsorption removal of dye Bromophenol blue from aqueous solution. The amounts of equilibrium adsorption were measured. The batch contact time method was used to measure the adsorption rate. Kinetic parameters were then evaluated and were finally correlated with process variables such as dye concentration, solution pH, contact time and temperature.

## MATERIALS AND METHODS

**Adsorbents and dye:** *Citrus reticulata* (orange) peels were obtained from a local fruit stall at Tirunelveli, India on November, 2009. The peels were cut into small pieces, crushed and washed thoroughly with deionized water to remove the adhering dirt. They were air dried in an oven at 40-50°C for 48 h. The BET surface areas of orange peels were in the range 0.0263-0.6719 m<sup>2</sup>/g obtained from N<sub>2</sub> adsorption isotherms by sorptiometer (Quantochrome NOVA 1000). The FR-IR spectra were obtained on Neclit 6700 model. Bromophenol blue was obtained from Merck Co. The solution pH was adjusted by adding a small amount of 0.1 M HCl or NaOH.

**Adsorption studies-Batch equilibrium experiments:** Bromophenol blue dye solutions were prepared by dissolving dye in deionized water to the required concentrations. In experiments of equilibrium adsorption isotherm, the mixture of dried adsorbent (containing 0.1 g dry basis of dye solution (50 cm<sup>3</sup>) and acetic acid buffer solution with desired pH value were shaken for 24 h using a water bath to control the temperature at 30±1°C. The effect of the adsorption isotherm was determined by examining the series of isotherms at different temperatures (30, 45 and 60°C) and for pH (5.5, 6.8, 7.8), respectively. The concentrations of dyes were measured with an UV/Visible spectrometer (JASCOV-530). The amount of dye absorbed onto the peels, q<sub>e</sub> (mg g<sup>-1</sup>), was calculated by a mass balance relationship (Eq. 1). The procedures of kinetic experiments were basically identical to those of equilibrium tests. The aqueous samples were taken at preset time intervals and the concentrations of dyes were similarly measured.

## RESULTS AND DISCUSSION

**Langmuir and freundlich isotherm:** The equilibrium adsorption isotherm is of fundamental importance in the design of adsorption systems. The isotherm expresses the relation between the mass of dye adsorbed at a particular temperature and pH and liquid phase of dye concentration. For any adsorption investigation one of the most important parameters required to understand the behavior of the adsorption process in the adsorption isotherm. The shape of an isotherm not only provides information about the affinity of the dye molecule for adsorption, but it also reflects the

possible mode of adsorbing dye molecule. The most common way of obtaining an adsorption isotherm, is to determine the concentration of dye solution before and after the adsorption experiments, although several attempts have been made to find the adsorbed amount. A basic assumption of the Langmuir theory (Langmuir, 1918) is that sorption takes place at specific sites within the adsorbent (Chen *et al.*, 2008; Asfour *et al.*, 1985; Poots *et al.*, 1976). The data obtained from the adsorption experiment conducted in the present investigation was fitted in different temperature and pH in isotherm equation as shown in Fig. 1A and B. The saturation monolayer can be represented by the expression.

$$q_e = \frac{KbC_e}{(1 + bC_e)} \quad (1)$$

$$\frac{1}{q_e} = \frac{1}{K} + \frac{1}{KbC_e} \quad (2)$$

A plot of  $(1/q_e$  vs  $1/C_e)$  resulted in a linear graphical relation indicating the applicability of the above model as shown in (Fig. 2A, B). The values are calculated from the slope and intercept of

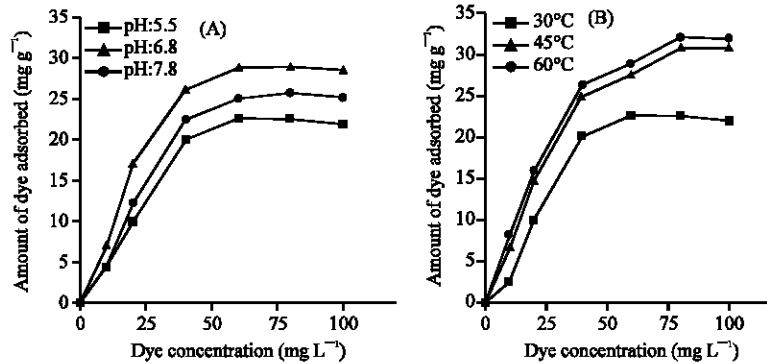


Fig. 1: Effect of specific dye uptakes (A) at different pH with dye concentration; (B) at different temperatures dye concentration

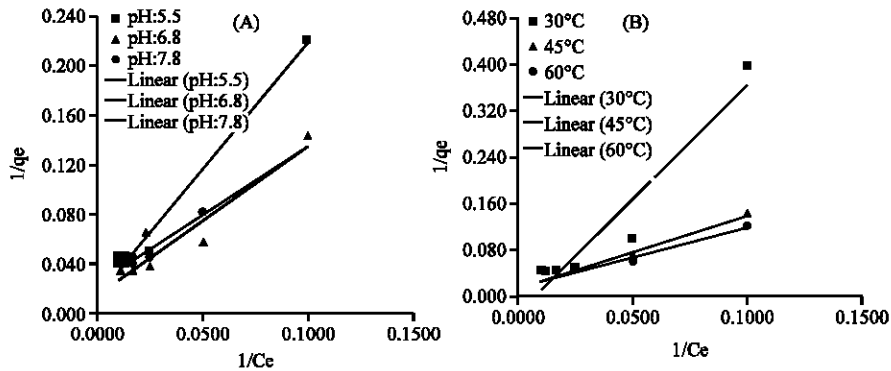


Fig. 2: Langmuir isotherm for the adsorption of dye using orange peels (A) at different pH with dye concentration; (B) at different temperature with dye concentration

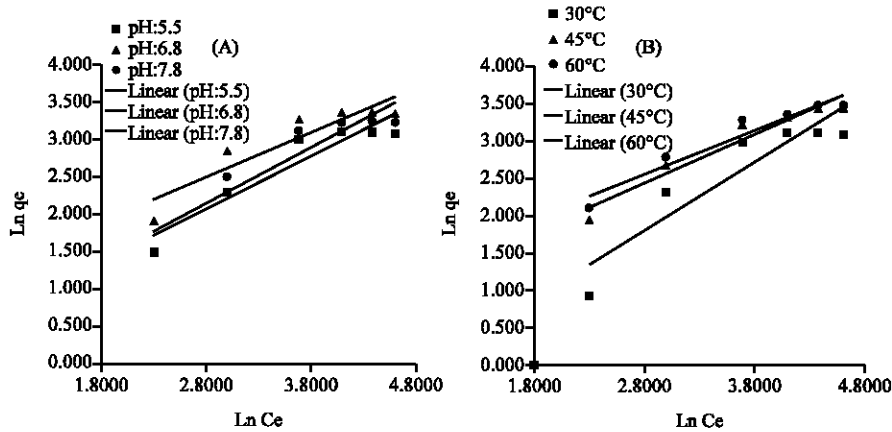


Fig. 3: Freundlich isotherm for the adsorption of dye using orange peels (A) at different pH with dye concentration and (B) at different temperature with dye concentration

different straight line representing the different pH and temperature (b) energy of adsorption and (k) adsorption capacity and  $Q_0$  is represented by (K). The Langmuir isotherm constant ( $Q_0$ ) in eqn (1) is a measure of the amount of dye adsorbed, when the monolayer is completed. Monolayer capacity ( $Q_0$ ) of the adsorbent for the dye is comparable obtained from adsorption isotherm. The observed statistically significant (at the 95% confidence level) linear relationship as evidenced of these by the  $R^2$  values (close to unity) indicate the applicability of the isotherm (Langmuir isotherm) and surface.

Freundlich isotherm (Freundlich, 1906) is used for heterogeneous surface energies system. The sorption isotherm is the most convenient form of representing the experimental data at different pH and temperature as shown in Fig. 3A and B. Figure 3A and B show the batch isothermal data fitted to the linear form of the Freundlich isotherm (Poots *et al.*, 1976; Ho and McKay, 1998, 1999; Mckay *et al.*, 1987; Namasivayam *et al.*, 1998):

$$q_e = K_F C_e^{1/n} \quad (3)$$

$$\ln q_e = \ln K_F + (1/n) \ln C_e \quad (4)$$

The various constants, associated with the isotherm are the intercept, which is roughly on indicator of sorption capacity ( $k_p$ ) and the slope ( $1/n$ ) sorption intensity values are recorded in (Table 1). Freundlich of isotherm has been illustrated to be a special case of heterogeneous surface energies and it can be easily extended to this case. It has been stated by Poots *et al.* (1976), Ho and McKay (1998, 1999), Mckay *et al.* (1987), Namasivayam *et al.* (1998) that magnitude of the exponent  $1/n$  gives an indication of the favorability and capacity of the adsorbent/adsorbate system. The values  $n > 1$  represents favorable adsorption conditions. Most of the cases the exponent between  $1 < n < 10$  shows the beneficial adsorption.

**Adsorption kinetics:** The kinetics of adsorption is important from the point of view that it controls the process efficiency. Various kinetic models have been used by various workers and different systems conform to different models but the Lagergrens rate equations (Lagergren, 1898) for the

Table 1: Langmuir and freundlich isotherm constants at different pH and temperature

Parameter	Langmuir Isotherm -model parameters	Freundlich Isotherm -model parameters
<b>pH</b>		
5.5	$K = 149.76; b = 0.49; R^2 = 0.9838$	$K_F = 0.093; n = 1.416; R^2 = 0.9433$
6.8	$K = 86.91; b = 0.82; R^2 = 0.9756$	$K_F = 0.8456; n = 1.769; R^2 = 0.8585$
7.8	$K = 25.74; b = 0.90; R^2 = 0.9359$	$K_F = 0.654; n = 1.344; R^2 = 0.8754$
<b>Temperature (°C)</b>		
30.0	$K = 86.91; b = 0.82; R^2 = 0.9756$	$K_F = 0.093; n = 1.416; R^2 = 0.9433$
45.0	$K = 88.11; b = 0.80; R^2 = 0.9822$	$K_F = 0.644; n = 1.55; R^2 = 0.9283$
60.0	$K = 62.72; b = 0.98; R^2 = 0.9879$	$K_F = 0.592; n = 1.68; R^2 = 0.9283$

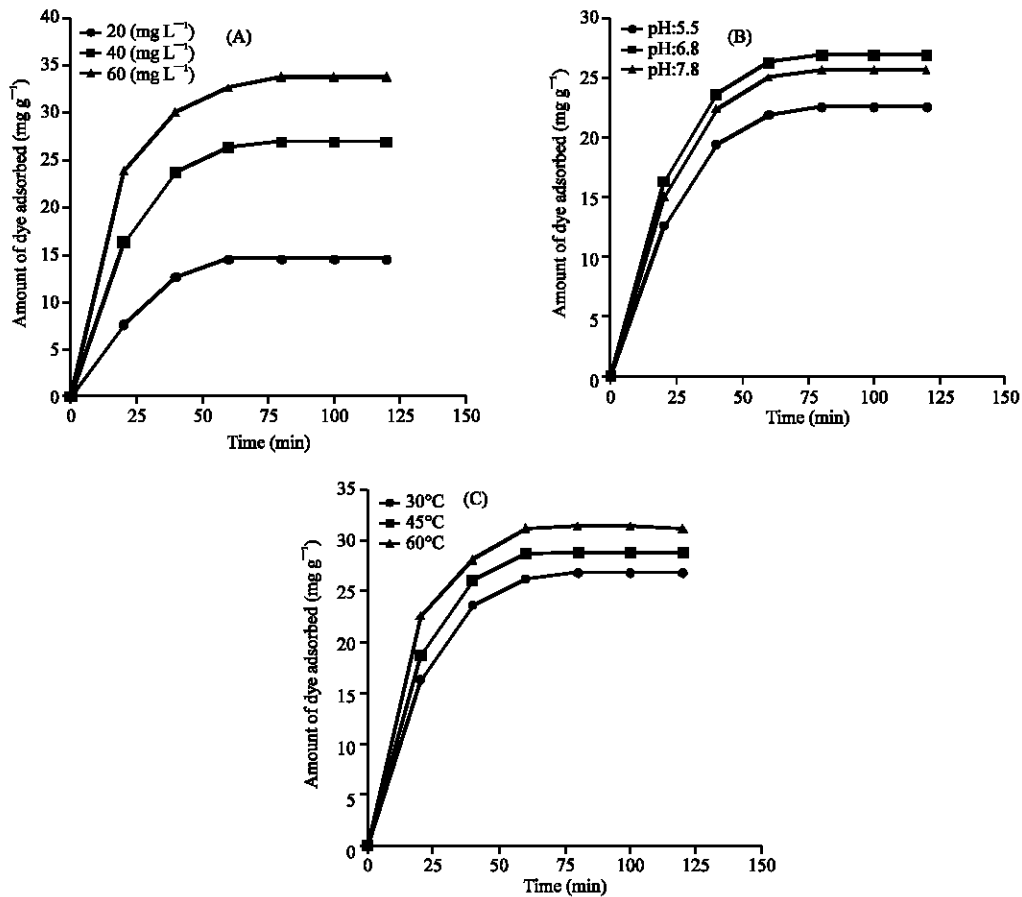


Fig. 4: Effect of specific dye uptakes (A) at different dye concentration with time (mints); (B) at different pH with time (mints) and (C) at different temperature with time (mints)

sorption of a solute form liquid solution. In order to examine the controlling mechanism of adsorption processes such as mass transfer and chemical reaction, several kinetic models are used to test experimental data (Namasivayam *et al.*, 2001; Namasivayam and Kavitha, 2002). Figure 4A-C had shown the adsorption of dye at different dye concentration, pH and temperature by using environmental nanomaterial. This may be due to the number of positive charges on the sorbent surface which leads to the no rejection of the negatively charged dye molecule and thereby increasing the adsorption. In general, the uptakes are much higher in acidic solutions than those in neutral and alkaline conditions. This explanation agrees with our data on pH effect. It can be

seen that the pH of aqueous solution plays an important role in the adsorption of Bromophenol blue dye onto Environmental Nanomaterial. This result is contrary to the results of Habib *et al.* (2006). Temperature is an important parameter for the adsorption process. A plot of the Bromophenol blue uptake as a function of temperature (30, 45 and 60°C) is shown in Fig. 4C. The adsorption of dye at higher temperature was found to be greater compared to that at a lower temperature. The curves indicate the strong tendency of the process for monolayer formation (Lucarelli *et al.*, 2000; Poots *et al.*, 1976; Ho and McKay, 1998, 1999; McKay *et al.*, 1987; Namasivayam *et al.*, 1998, 2001). The increase in temperature would increase the mobility of the large dye ion and also produces a swelling effect with in the internal structure of the environmental nanomaterial, thus enabling the large dye molecule to penetrate further (Namasivayam *et al.*, 2001; Namasivayam and Kavitha, 2002; Ho and McKay, 1999). Therefore, the adsorption capacity should largely depend on the chemical interaction between the functional groups on the adsorbent surface and the adsorbate and should increase with temperature rising. The adsorption of dye at higher temperature was found in the present investigation is similar to results of Namasivayam *et al.* (2001), Namasivayam and Kavitha (2002) and Ho and McKay (1999).

**Pseudo-first and second-order equation:** In system design viewpoint, a lumped analysis of adsorption rates is thus sufficient to practical operation. A simple kinetic analysis of adsorption is the pseudo-first-order equation:

$$\frac{dq_t}{q_t} = K_1 (q_{eq} - q_t) \quad (5)$$

After definite integration by applying the initial conditions  $q_t=0$  at  $t=0$  and  $q_t=q_t$  at  $t=t$ , Eq. 4 becomes:

$$\log(q_{eq} - q_t) = \log q_{eq} - \frac{K_1}{2.303} t \quad (6)$$

where,  $q_{eq}$  and  $q_t$  are amount of dye adsorbed at equilibrium and at time, in mg g<sup>-1</sup>, respectively and  $K_1$  is the first order rate constant, was applied to the present studies of dye adsorption. As such the values of  $\log (q_{eq} - q_t)$  vs  $t$  were calculated from the kinetic data of Fig. 5A-C and plotted against time. The first-order rate constant calculated from the plots are shown in Table 2. Adsorption kinetics for some system can also be described by a pseudo-second order reaction (Ho and McKay, 1998, 1999; McKay *et al.*, 1987; Namasivayam *et al.*, 1998, 2001; Namasivayam and Kavitha, 2002).

The pseudo-second-order equation based on adsorption equilibrium capacity may be expressed in the form:

$$\frac{dq_t}{q_t} = K_2 (q_{eq} - q_t)^2 \quad (7)$$

where,  $k_2$  is the rate constant of pseudo-second-order adsorption. Integrating Eq. 6 and applying the initial conditions, we have

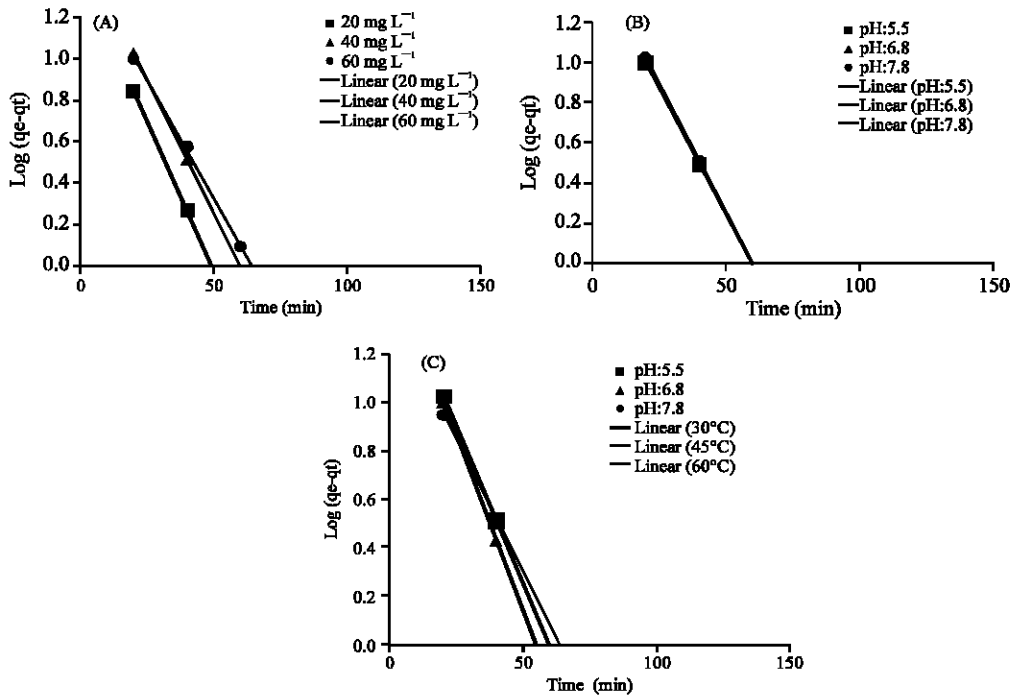


Fig. 5: Pseudo-first order plot for the adsorption of dye using orange peels (A) at various plots at dye concentration; (B) at various plots at pH and (C) at various plots at temperatures

Table 2: Pseudo-first and Pseudo-Second order rate constant at different dye concentration, pH and temperature

Parameter	Pseudo-first order rate constant	Pseudo-Second order rate constant
<b>Dye concentration (mg L<sup>-1</sup>)</b>		
20.0	$K_1 = 0.0658; q_{e,q} = 1.4165; R^2 = 0.99989$	$K_2 = 0.0442; q_{e,q} = 14.51; R^2 = 0.9989$
40.0	$K_1 = 0.258; q_{e,q} = 1.5401; R^2 = 0.9999$	$K_2 = 0.0075; q_{e,q} = 34.96; R^2 = 0.9986$
60.0	$K_1 = 0.226; q_{e,q} = 1.4597; R^2 = 0.9989$	$K_2 = 0.0156; q_{e,q} = 27.47; R^2 = 0.9987$
<b>pH</b>		
5.5	$K_1 = 0.0258; q_{e,q} = 1.4165; R^2 = 0.9989$	$K_2 = 0.01545; q_{e,q} = 23.148; R^2 = 0.9995$
6.8	$K_1 = 0.0257; q_{e,q} = 1.5401; R^2 = 0.9999$	$K_2 = 0.01555; q_{e,q} = 26.246; R^2 = 0.9996$
7.8	$K_1 = 0.0258; q_{e,q} = 1.4597; R^2 = 0.9896$	$K_2 = 0.01566; q_{e,q} = 27.47; R^2 = 0.9997$
<b>Temperature (°C)</b>		
30.0	$K_1 = 0.0257; q_{e,q} = 1.4165; R^2 = 0.9989$	$K_2 = 0.10935; q_{e,q} = 14.90; R^2 = 0.9995$
45.0	$K_1 = 0.0287; q_{e,q} = 1.5401; R^2 = 0.9998$	$K_2 = 0.0038; q_{e,q} = 25.90; R^2 = 0.9686^5$
60.0	$K_1 = 0.0218; q_{e,q} = 1.4597; R^2 = 0.9989$	$K_2 = 0.0476; q_{e,q} = 21.08; R^2 = 0.9935$

$$\frac{1}{(q_{eq} - q_t)} = \frac{1}{q_e} + K_2 t \tag{8}$$

or equivalently,

$$\frac{t}{q_t} = \frac{1}{K_2 q_{eq}^2} + \frac{1}{q_e} t \tag{9}$$

The equilibrium adsorption capacity ( $q_{e,q}$ ) and the second-order constants  $k_2$  (g/mg min) can be determined experimentally from the slope and intercept of plot  $t/q_t$  versus  $t$ . The applicability of the

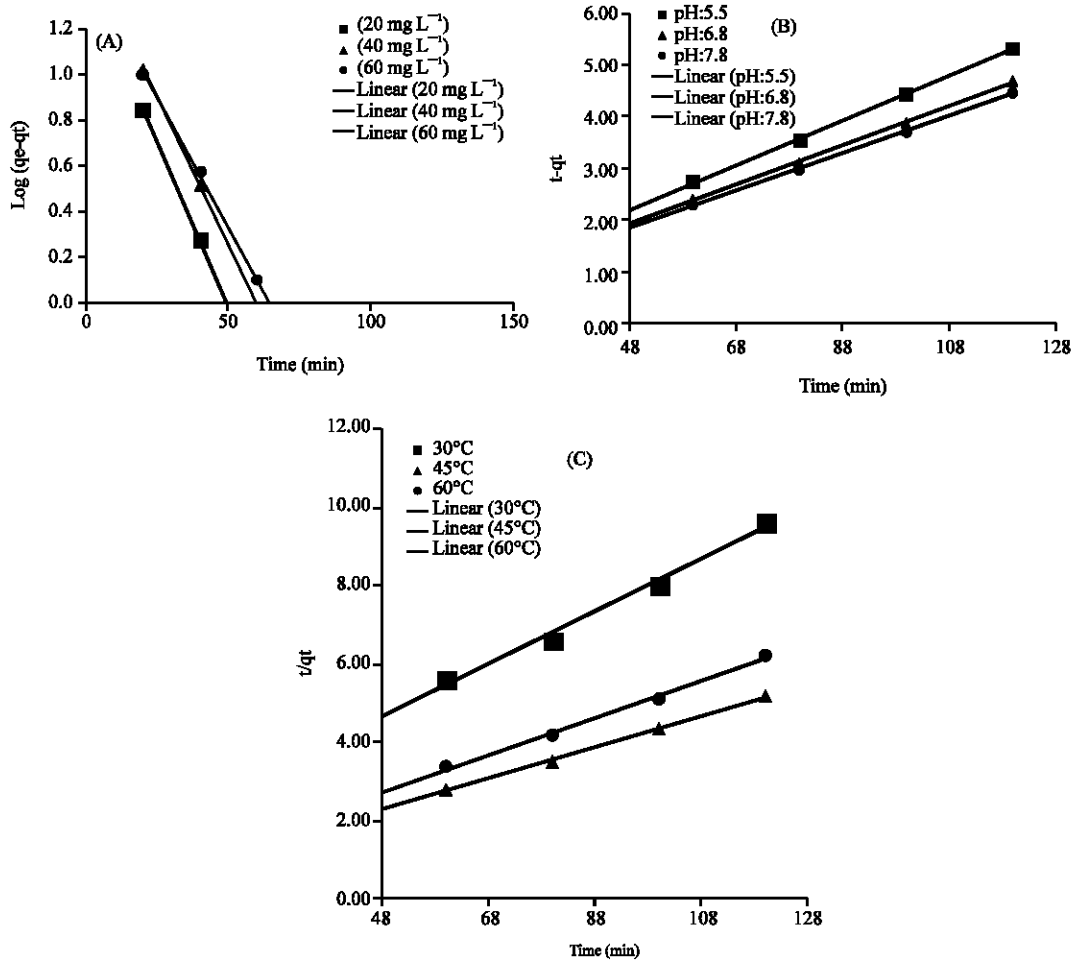


Fig. 6: Pseudo-second order plot for the adsorption of dye using orange peels (A) at various plots at dye concentration; (B) at various plots at pH and (C) at various plots at temperatures

pseudo-second order models can be examined by linear plot  $t/q$  vs  $t$  respectively as shown in Fig. 6A-C. The correlation coefficient  $R^2$  shows that the pseudo-second order model an indications of a chemisorptions mechanism, fits the experimental data slightly better than the pseudo-first order model. Therefore the adsorption of Bromophenol blue dye can be approximated more favorably by the pseudo-second order model. This model has been successfully applied to describe the kinetics of many adsorption systems. Calculated correlations are closer to unity for second-order kinetics model; therefore the adsorption kinetics could well be approximated more favorably by second-order kinetic model for dye adsorption. The  $k_2$  ( $\text{mg g}^{-1} \text{min}^{-1}$ ) and  $q_{eq}$  ( $\text{mg g}^{-1} \text{min}^{-1}$ ) values as calculated are listed in Table 2.

**Characterization of orange fruit peels:** The constituents of these peels were determined and characterized by FT-IR as shown in Fig. 7. The FT-IR spectra of these peels indicated that these adsorbents contained carboxyl groups between  $1257\text{-}1244 \text{ cm}^{-1}$  and hydroxyl groups at  $3423 \text{ cm}^{-1}$



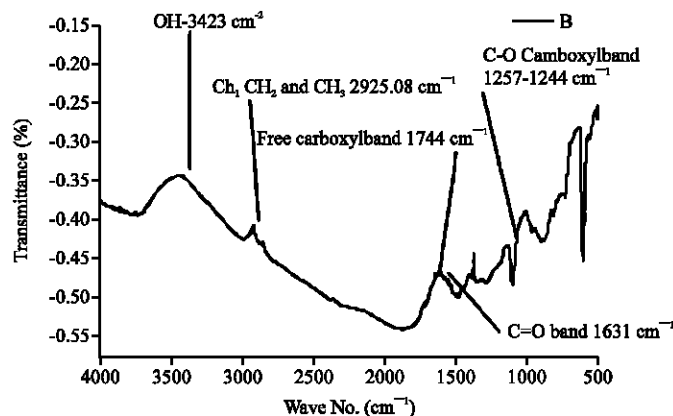


Fig. 7: Orange fruit peel cellulose FT-IR Spectrum

were determined. It clearly revealed that the carboxyl and hydroxyl groups content will influence the adsorption capacities and higher the group's content, the better was the adsorption capacities.

## CONCLUSION

The adsorption of dye from aqueous solution using orange cellulose-based waste peels has been investigated, under different reaction conditions in batch and equilibrium mode. The fitness of Langmuir model in the present system shows the formation of monolayer coverage of the adsorbate at the outer space of the adsorbent. Freundlich model isotherm was analyzed. The monolayer adsorption capacity determined was reasonably high ( $\text{mg g}^{-1}$ ) at  $n$ , Temperature (30, 45, 60°C) 88.11 ( $\text{mg g}^{-1}$ ) and pH 149.760 ( $\text{mg g}^{-1}$ ) for adsorption of Bromophenol blue dye, respectively. The values of dimensionless equilibrium parameter like separation factor ( $R_L$ ) at different particle size and temperature indicates the favorability of the process described in the present study. The kinetics of Bromophenol blue dye adsorption nicely followed pseudo-first and second order rate expression and a significant role in the adsorption mechanism. Langmuir and Freundlich models could be used to describe dye sorption on environmental nanomaterial at equilibrium and the kinetic data also gave a better fit. The data reported here should be useful for the design and fabrication of an economically viable treatment process using batch (or) stirred tank reactors and also it's revealed that the waste orange peels were the potential materials for dyes removal from aqueous solutions.

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